

Acids, Bases, and Solutions ▪ *Reading/Notetaking Guide***Concentration and Solubility** (pp. 262–267)

This section describes how concentration is measured. It also describes the usefulness of solubility and factors that affect it.

Use Target Reading Skills

After you read the section, for each Key Term write a meaningful sentence that incorporates that Key Term.

dilute solution

concentrated solution

solubility

saturated solution

unsaturated solution

supersaturated solution

Concentration (pp. 262–263)

Match the term with its definition.

Term

- ____ 1. dilute solution
____ 2. concentrated solution

Definition

- a. A mixture that has a lot of solute dissolved in it
b. A mixture that has only a little solute dissolved in it

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3. What are two ways in which you can change the concentration of a solution?

4. How do you measure the concentration of a solution?

Solubility (pp. 263–264)

5. What is solubility?

6. A mixture that has so much solute in it that no more will dissolve is called a(n)

_____.

7. A mixture in which more solute can be dissolved is called a(n)

_____.

8. Which is more soluble in water, baking soda or sugar? _____

9. Is the following sentence true or false? Solubility can be used to identify an unknown substance. _____

Factors Affecting Solubility (pp. 264–267)

10. What are three factors that affect the solubility of a substance?

a. _____ b. _____
c. _____

11. The higher the pressure of the gas, the _____ gas can dissolve in a solvent.

12. Is the following sentence true or false? Nonpolar compounds usually dissolve in polar solvents. _____

13. Circle the letter of each sentence that is true about temperature and solubility.

- a. Most solids become more soluble as the temperature goes up.
- b. Most gases become less soluble as the temperature goes up.
- c. Sugar dissolves better in cold water than in hot water.
- d. Carbon dioxide dissolves better in cold water than in hot water.