

The Nature of Matter ▪ *Reading/Notetaking Guide*

Describing Matter (pp. 58–67)

This section describes the kinds of properties used to describe matter. It also defines elements and contrasts compounds and mixtures.

Use Target Reading Skills

Write a definition of each Key Term in your own words.

matter: _____

chemistry: _____

substance: _____

physical property: _____

chemical property: _____

element: _____

atom: _____

chemical bond: _____

molecule: _____

compound: _____

chemical formula: _____

mixture: _____

heterogeneous mixture: _____

homogeneous mixture: _____

solution: _____

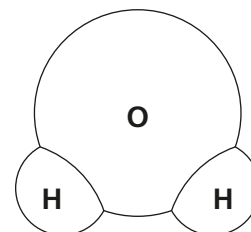
The Nature of Matter ▪ *Reading/Notetaking Guide***Properties of Matter** (pp. 59–61)

1. The study of the properties of matter and how matter changes is called _____.
2. Is the following sentence true or false? Table sugar and table salt are pure substances. _____
3. A(n) _____ property is a characteristic of a pure substance that can be observed without changing the substance into something else.
4. Complete the table by classifying each property as either a physical or chemical property.

Properties of Matter	
Property	Physical or Chemical?
Ability to burn	a.
Color	b.
Flexibility	c.
Ability to tarnish	d.
Melting point	e.
Ability to rust	f.

Elements (pp. 62–63)

5. A pure substance that cannot be broken down into any other substances by chemical or physical means is a(n) _____.
6. Is the following sentence true or false? The basic particle from which all elements are made is a molecule. _____
7. When atoms combine, the force of attraction that holds them together is a(n) _____.
8. The diagram to the right represents a water molecule. How many atoms of hydrogen does a water molecule contain? _____



The Nature of Matter ▪ *Reading/Notetaking Guide*

Describing Matter *(continued)*

Compounds (p. 64)

9. What is a compound?

10. What is the ratio of atoms in carbon dioxide, or CO₂?

11. What is the chemical formula of carbon monoxide?

12. Is the following sentence true or false? When elements chemically combine, they form compounds that have properties that are similar to those of the uncombined elements. _____

Mixtures (pp. 65–67)

13. A(n) _____ is made of two or more substances that are together in the same place but are not chemically combined.

14. What are two ways in which mixtures differ from compounds?

a. _____

b. _____

15. Circle the letter of each mixture below that is heterogeneous.

- a. damp soil
- b. sugar water
- c. brass
- d. salad

16. Is the following sentence true or false? A solution is an example of a homogeneous mixture. _____

17. Give three examples of ways to separate the substances in a mixture.
