

Acids, Bases, and Solutions ▪ *Chapter 7 Pre-Assessment*

Write the letter of the correct answer on the line at the left.

- _____ 1. A solution is an example of a
 - a. homogeneous colloid.
 - b. heterogeneous colloid.
 - c. homogeneous mixture.
 - d. heterogeneous mixture.
- _____ 2. Magnesium sulfide and aluminum fluoride are
 - a. ionic compounds.
 - b. molecular compounds.
 - c. covalent electrons.
 - d. radioactive elements.
- _____ 3. When dissolved in water, ionic compounds
 - a. conduct electricity.
 - b. make the water cloudy.
 - c. form double and triple bonds.
 - d. do not conduct electricity.
- _____ 4. When dissolved in water, molecular compounds
 - a. conduct electricity.
 - b. make the water cloudy.
 - c. form double and triple bonds.
 - d. do not conduct electricity.

Acids, Bases, and Solutions ▪ *Section 7.1 Quiz*

If the statement is true, write true. If it is false, change the underlined word or words to make the statement true.

- _____ 1. Sodium chloride and other compounds are solvents in salt water.
- _____ 2. Water dissolves so many substances that it is often called the "universal solvent."
- _____ 3. Solutions can be formed from any combination of liquids.
- _____ 4. Colloids and suspensions have different properties than solutions.
- _____ 5. Salt lowers the boiling point and freezing point of water.

Acids, Bases, and Solutions ▪ Section 7.2 Quiz

Fill in the blank to complete each statement.

1. You can make a concentrated solution by adding more _____ or removing _____.
2. You can make a dilute solution by increasing the amount of _____ in a solution.
3. To measure concentration, you compare the amount of _____ to the total amount of solution.
4. You can identify a substance by its solubility because solubility is a _____ property of matter.
5. Three factors that affect the solubility of a substance are _____, _____, and _____.

Acids, Bases, and Solutions ▪ Section 7.3 Quiz

If the statement is true, write true. If it is false, change the underlined word or words to make the statement true.

- _____ 1. Acids react with certain metals to produce carbon dioxide gas.
- _____ 2. Acids react in a characteristic way with cabornate ions, producing carbon dioxide gas.
- _____ 3. Common acids include sodium hydroxide, calcium hydroxide, and ammonia.
- _____ 4. The reactions of the base baking soda with an acid, such as buttermilk, makes biscuits light and fluffy.
- _____ 5. You can find acids and bases exclusively in laboratories and research institutions.

Acids, Bases, and Solutions ▪ *Section 7.4 Quiz*

Fill in the blank to complete each statement.

1. _____ ions are the key to the reactions of acids.
2. In a solution of a _____ acid, all the acid molecules break up into ions.
3. Knowing the _____ is the key to knowing how acidic or basic a solution is.
4. The _____ expresses the concentration of hydrogen ions in a solution.
5. In a neutralization reaction, an acid reacts with a base to produce _____.